



# **Mark Scheme (Results)**

**Summer 2018**

Pearson Edexcel GCE  
In Chemistry (8CH0) Paper 01  
Core Inorganic and Physical Chemistry

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Summer 2018

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## General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be **prepared to award zero marks if the candidate's response is not** worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the **mark scheme to a candidate's response, the team leader must** be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.

## Using the Mark Scheme

Examiners should look for qualities to reward rather than faults to penalise. This does NOT mean giving credit for incorrect or inadequate answers, but it does mean allowing candidates to be rewarded for answers showing correct application of principles and knowledge. Examiners should therefore read carefully and consider every response: even if it is not what is expected it may be worthy of credit.

The mark scheme gives examiners:

- an idea of the types of response expected
- how individual marks are to be awarded
- the total mark for each question
- examples of responses that should NOT receive credit.

/ means that the responses are alternatives and either answer should receive full credit.

( ) means that a phrase/word is not essential for the award of the mark, but helps the examiner to get the sense of the expected answer.

Phrases/words in **bold** indicate that the meaning of the phrase or the actual word is essential to the answer.

ecf/TE/cq (error carried forward) means that a wrong answer given in an earlier part of a question is used correctly in answer to a later part of the same question.

Candidates must make their meaning clear to the examiner to gain the mark. Make sure that the answer makes sense. Do not give credit for correct words/phrases which are put together in a meaningless manner. Answers must be in the correct context.

### Quality of Written Communication


Questions which involve the writing of continuous prose will expect candidates to:

- write legibly, with accurate use of spelling, grammar and punctuation in order to make the meaning clear
- select and use a form and style of writing appropriate to purpose and to complex subject matter
- organise information clearly and coherently, using specialist vocabulary when appropriate.

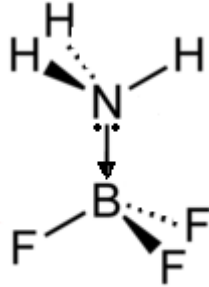
Full marks will be awarded if the candidate has demonstrated the above abilities.

Questions where QWC is likely to be particularly important are indicated (QWC) in the mark scheme, but this does not preclude others.

| Question Number | Acceptable Answer  | Additional Guidance | Mark |
|-----------------|--|---------------------|------|
| 1(a)            | <p>An answer that makes reference to the following points:</p> <ul style="list-style-type: none"> <li>(strong electrostatic) attraction (1)</li> <li>between two nuclei and the shared /bonding pair of electrons (1)</li> </ul> |                     | (2)  |

| Question Number | Acceptable Answer   | Additional Guidance  | Mark |
|-----------------|---|--|------|
| 1(b)            | <ul style="list-style-type: none"> <li>diagram showing 3-D shape of ammonia, including two bonds <b>with one 'wedge' and one 'hatch' and one N-H bond 'in plane'</b> (1)</li> <li>lone pair of electrons on nitrogen atom and</li> <li>bond angle of <math>107^\circ</math> labelled (1)</li> </ul> | <p><u>Example of diagram:</u></p>  <p>Allow any direction of the wedge and/or hatch</p> <p>This mark can be scored on a dot and cross diagram</p> <p>Allow any angle between <math>106^\circ</math> and <math>108^\circ</math> inclusive.</p> <p>Do not award M2 if the <math>107^\circ</math> bond angle is shown as that between the lone pair and a bonding pair</p> <p>Ignore name of shape even if incorrect</p> | (2)  |

| Question Number | Answer  | Mark |
|-----------------|---|------|
| 1(c)            | <p>The only correct answer is D</p> <p><i>A is not correct because this is approximately the angle given in the diagram</i></p> <p><i>B is not correct because this is the angle for three bonds when there is also a lone pair on the central atom</i></p> <p><i>C is not correct because this is the angle when there are four pairs of bonding electrons around the central atom</i></p> | (1)  |

| Question Number | Acceptable Answer   | Additional Guidance   | Mark |
|-----------------|---|---|------|
| 1(d)(i)         | <p>An answer that makes reference to the following points:</p> <ul style="list-style-type: none"> <li>• donation of lone pair (of electrons) from nitrogen / lone pair from ammonia (1)</li> <li>• to the boron (atom) which is electron deficient / has only 6 electrons in outer shell / has 6 valence electrons / can accept two electrons to complete octet / can accept two electrons to get a full (outer) shell (1)</li> </ul> | <p><b>Allow 'non-bonding pair' for lone pair</b><br/> <b>Allow 'sharing' for donation</b></p> <p>Do not penalise donation to F atoms, but can only score M1 in this case</p> <p>Allow <b>just</b> 'boron has an incomplete outer shell'<br/> Allow boron has an empty (p-)orbital</p> <p>Do not award M2 for <b>just</b> 'nitrogen shares lone pair with boron atom' or similar</p> <p>M1 may be scored from a diagram here OR a diagram in (d)(ii)<br/> e.g.</p>  <p>scores only M1</p> | (2)  |

| Question Number | Acceptable Answer  | Additional Guidance  | Mark |
|-----------------|--|--|------|
| 1(d)(ii)        | <ul style="list-style-type: none"> <li data-bbox="383 443 1249 480">• HNH angle is (approximately) <math>109.5^\circ</math> (1)</li> <br/> <li data-bbox="383 587 1249 624">• FBF angle is (approximately) <math>109.5^\circ</math> (1)</li> </ul> | <p data-bbox="1301 411 1877 517">May be shown on a diagram, including on a diagram in 3(d)(i) e.g</p> <div data-bbox="1496 549 1720 847" style="text-align: center;"> </div> <p data-bbox="1301 935 1818 1040">Allow 1 for just <math>109.5^\circ</math> if it has not been made clear that this angle applies to BOTH bond angles</p> <p data-bbox="1301 1082 1742 1150"><u>Both angles</u> change to <math>109.5^\circ</math> scores 2</p> <p data-bbox="1301 1192 1532 1219">Allow 109-110°</p> | (2)  |

(Total for Question 1 = 9 marks)



| Question Number | Acceptable Answer  | Additional Guidance     | Mark |
|-----------------|--|-------------------------|------|
| 2(a)(i)         | $({}^1_1\text{H})$ protons 1, neutrons 0<br>$({}^2_1\text{H})$ protons 1, neutrons 1 | All four correct needed | (1)  |

| Question Number | Acceptable Answer   | Additional Guidance                | Mark |
|-----------------|---|------------------------------------|------|
| 2(a)(ii)        | An explanation that makes reference to the following points: <ul style="list-style-type: none"> <li>• (atoms that) have the same number of protons (1)</li> <li>• but a different number of neutrons (1)</li> </ul> | Ignore any references to electrons | (2)  |

| Question Number | Acceptable Answer  | Additional Guidance   | Mark |
|-----------------|--|---|------|
| 2(b)(i)         | An answer that makes reference to following: <ul style="list-style-type: none"> <li>• both isotopes have an isotopic mass of greater than 1 / 1.0 / one</li> </ul> OR <ul style="list-style-type: none"> <li>• there are no isotopes with an isotopic mass of less than one</li> </ul> | Award mark if it is stated that the (only) other isotope is ${}^2\text{H}$<br><br>Ignore calculation of value, even if incorrect. | (1)  |

| Question Number | Acceptable Answer   | Additional Guidance   | Mark |
|-----------------|---|---|------|
| 2(b)(ii)        | <ul style="list-style-type: none"> <li>calculation to find <math>A_r</math> (1)</li> <li>value of <math>A_r</math> to 4 DP (1)</li> </ul> | <p><u>Example of calculation</u></p> $A_r = \frac{(1.007825 \times 99.9885) + (2.014101 \times 0.0115)}{100}$ <p>(= 1.0079407) = 1.0079<br/> Correct answer with no working scores (2)</p> <p>Allow TE for M2 for incorrect transfer of data or for one incorrect % abundance (e.g. 1.15%), provided that the final <math>A_r</math> value is between 1 and 2</p> <p>Ignore units even if incorrect</p> | (2)  |

| Question Number | Acceptable Answer  | Additional Guidance   | Mark |
|-----------------|--|---|------|
| 2(c)(i)         | <p>An answer that makes reference to the following:</p> <ul style="list-style-type: none"> <li>equation (1)</li> <li>state symbol, (g), on both H and H<sup>+</sup> (1)</li> </ul> | <p>H(g) → H<sup>+</sup>(g) + e<sup>(-)</sup></p> <p>or</p> <p>H(g) - e<sup>(-)</sup> → H<sup>+</sup>(g)</p> <p>Ignore state symbol for electron</p> <p>H<sub>2</sub>(g) → H<sub>2</sub><sup>+</sup>(g) + e<sup>(-)</sup> scores only M2<br/> H<sub>2</sub>(g) - e<sup>(-)</sup> → H<sub>2</sub><sup>+</sup>(g) scores only M2<br/> H<sub>2</sub>(g) → 2H<sup>+</sup>(g) + 2e<sup>(-)</sup> scores 0<br/> X(g) → X<sup>+</sup>(g) + e<sup>(-)</sup> scores only M2</p> | (2)  |

| Question Number | Acceptable Answer  | Additional Guidance   | Mark |
|-----------------|--|---|------|
| 2(c)(ii)        | <p>An explanation that makes reference to the following points:</p> <p><u>H &lt; He:</u></p> <ul style="list-style-type: none"> <li>• He more protons than H / He greater nuclear charge than H (1)</li> <li>• in helium the outer electron is in the same shell as hydrogen (1)</li> </ul> <p><u>H &gt; Li:</u></p> <ul style="list-style-type: none"> <li>• in lithium the outer electron is in a higher energy level / a new shell / further from the nucleus / in a 2s orbital (1)</li> <li>• (and) is shielded by inner electrons / 1s<sup>2</sup> electrons (1)</li> </ul> | <p>Ignore references to shielding for H and He</p> <p>Ignore references to atomic radius or electrons being closer to or the same distance from the nucleus in helium</p> <p>Allow lithium has more shells of electrons</p> <p>Allow (outer electron of) lithium has more shielding than hydrogen / is shielded</p> | (4)  |

| Question Number | Acceptable Answer  | Additional Guidance  | Mark |
|-----------------|--|--|------|
| 2(d)            | <p>An answer that makes reference to the following:</p> <p>(in favour)</p> <ul style="list-style-type: none"> <li>• electronic structure of hydrogen is <math>s^1</math> / <math>1s^1</math> / has one electron in s orbital / form <math>1+</math> ions (1)</li> </ul> <p>(against) any two from</p> <ul style="list-style-type: none"> <li>• the rest of Group 1 are (alkali) metals / metallic (hydrogen is not) (1)</li> <li>• hydrogen does not react in the same way as / has different reactivity to the rest of Group 1 / has different chemical properties (1)</li> <li>• forms a <math>H^-</math> ion (1)</li> </ul> | <p>Allow 1 electron in outer shell / has 1 valence electron<br/> <b>Do not award 'last electron is in s orbital' unless</b> it is clear there is only one<br/> <b>Do not award just 'single unpaired electron'</b></p> <p>Allow hydrogen is not a metal<br/> Ignore hydrogen is a gas but Group 1 elements are solid</p> <p><b>Do not award just 'different properties' or 'different behaviour'</b><br/> Allow hydrogen forms covalent bonds as a chemically different property<br/> Ignore trends in physical properties</p> <p>Allow hydrogen can gain one electron to form a stable ion / become stable / fill its outer shell</p> | (3)  |

(Total for Question 2 = 15 marks)

| Question Number | Acceptable Answer | Mark |
|-----------------|-------------------|------|
|-----------------|-------------------|------|

|      |  |     |
|------|--|-----|
| 3(a) | <p>The only correct answer is C</p> <p><i>A is not correct because a burette is used to measure varied volumes</i></p> <p><i>B is not correct because a measuring cylinder is less precise</i></p> <p><i>D is not correct because a volumetric flask is less precise</i></p> | (1) |
|------|--|-----|

| Question Number | Acceptable Answer  | Mark |
|-----------------|--|------|
| 3(b)            | <p>The only correct answer is C</p> <p><i>A is not correct because this is the appearance of the solution before the potassium hydroxide is added</i></p> <p><i>B is not correct because this is the colour that methyl orange would be in neutral solution</i></p> <p><i>D is not correct because this is a colour sometimes given for the end-point which is incorrect, and it is the colour of phenolphthalein in acidic solution</i></p> | (1)  |

| Question Number | Acceptable Answer   | Additional Guidance   | Mark |
|-----------------|---|---|------|
| 3(c)(i)         | <ul style="list-style-type: none"> <li>two correct readings to nearest 0.05 (1)</li> <li>correct subtraction of two values to 2 d.p. (1)</li> </ul> | <p><u>Example of answer</u><br/>32.35 and 4.60</p> <p>27.75<br/>Allow TE for M2 on their burette readings</p> | (2)  |

| Question Number | Acceptable Answer  | Mark |
|-----------------|--|------|
| 3(c)(ii)        | <p>The only correct answer is A</p> <p><i>B is not correct because this is the mean of the three values given without the rough value</i></p> <p><i>C is not correct because this is the mean of the last two values</i></p> <p><i>D is not correct because this is the mean of all four including the rough value</i></p> | (1)  |

| Question Number | Acceptable Answer   | Additional Guidance  | Mark |
|-----------------|---|--|------|
| 3(c)(iii)       | <ul style="list-style-type: none"> <li>• calculates moles of H<sub>2</sub>SO<sub>4</sub> (1)</li> <li>• calculates moles of KOH (1)</li> <li>• calculates concentration of KOH to 2/3 SF (1)</li> </ul> | <p><u>Example of calculation</u></p> <p>= <math>0.0800 \times \frac{25}{1000} = 0.00200</math> (mol)</p> <p>= <math>0.00200 \times 2 = 0.00400</math> (mol)</p> <p>= <math>\frac{0.00400}{27.00} \times 1000 (= 0.148148148\dots)</math> (mol dm<sup>-3</sup>)</p> <p>= <math>0.148/0.15</math> (mol dm<sup>-3</sup>) to 2 or 3 SF</p> <p>Allow TE on all stages of the calculation</p> <p>Correct answer with no working scores (3)</p> | (3)  |

(Total for Question 3 = 8 marks)

| Question Number | Acceptable Answer  | Additional Guidance     | Mark |
|-----------------|--|-------------------------|------|
| 4(a)(i)         | <p>An answer that makes reference to two of the following:</p> <ul style="list-style-type: none"> <li>sulfate / sulfate(VI) / <math>\text{SO}_4^{2-}</math></li> <li>sulfite / sulfate(IV) / <math>\text{SO}_3^{2-}</math></li> <li>carbonate / <math>\text{CO}_3^{2-}</math></li> </ul> | Penalise lack of charge | (1)  |

| Question Number | Acceptable Answer  | Additional Guidance   | Mark |
|-----------------|--------------------|---|------|
| 4(a)(ii)        | $\text{SO}_4^{2-}$ | <p>Ignore sulfate (ion)</p> <p>Only penalise lack of charge if not penalised in 4(a)(i)</p> | (1)  |

| Question Number | Acceptable Answer  | Mark |
|-----------------|--|------|
| 4(a)(iii)       | <p>The only correct answer is C</p> <p><i>A is not correct because the ratio is one-to-one</i></p> <p><i>B is not correct because cations are positive</i></p> <p><i>D is not correct because cations are positive</i></p> | (1)  |

| Question Number | Acceptable Answer                            | Additional Guidance   | Mark |
|-----------------|--|---|------|
| 4(b)            | Cation is $\text{Mg}^{2+}$ / magnesium (ion) | <p><b>Do not award use of symbol just "Mg"</b></p> <p>Award <math>\text{Be}^{2+}</math> / beryllium (ion)</p> | (1)  |

| Question Number             | Acceptable Answer  | Additional Guidance  | Mark |          |        |   |      |      |                        |                                   |                                |                             |   |            |  |          |        |   |      |      |                        |                          |                      |                             |   |        |     |
|-----------------------------|--|--|------|----------|--------|---|------|------|------------------------|-----------------------------------|--------------------------------|-----------------------------|---|------------|--|----------|--------|---|------|------|------------------------|--------------------------|----------------------|-----------------------------|---|--------|-----|
| 4(c)                        | <p>Method 1<br/>Calculates <math>M_r</math> of <math>MgSO_4</math> (1)</p> <p>Divides percentage by relative formula mass (1)</p> <p>Divides ratio by smallest (1)</p> <p><math>x = 7</math> (1)</p> <p>For Alternative<br/>Calculates <math>M_r</math> of <math>CoCl_2</math> (1)</p> <p>Divides percentage by relative formula mass (1)</p> <p>Divides ratio by smallest (1)</p> <p><math>y = 6</math> (1)</p> | <p>In all cases correct answer with some correct working scores (4)</p> <p><u>Example of calculation:</u><br/> <math>M_r</math> of <math>MgSO_4 = 24.3 + 32.1 + (4 \times 16) = 120.4</math></p> <table border="1" data-bbox="1146 459 1874 746"> <thead> <tr> <th></th> <th><math>MgSO_4</math></th> <th><math>H_2O</math></th> </tr> </thead> <tbody> <tr> <td>%</td> <td>48.9</td> <td>51.1</td> </tr> <tr> <td>Moles<br/>(<math>\div</math> RFM)</td> <td><math>48.9 / 120.4 = 0.406146170\dots</math></td> <td><math>51.1 / 18 = 2.838888889\dots</math></td> </tr> <tr> <td>Ratio<br/>(<math>\div</math> smallest)</td> <td>1</td> <td>6.98982049</td> </tr> </tbody> </table> <p>Allow <math>MgSO_4 \cdot 7H_2O</math></p> <p><math>M_r</math> of <math>CoCl_2 = 58.9 + 2 \times 35.5 = 129.9</math></p> <table border="1" data-bbox="1146 997 1874 1284"> <thead> <tr> <th></th> <th><math>CoCl_2</math></th> <th><math>H_2O</math></th> </tr> </thead> <tbody> <tr> <td>%</td> <td>54.6</td> <td>45.4</td> </tr> <tr> <td>Moles<br/>(<math>\div</math> RFM)</td> <td><math>54.6 / 129.9 = 0.42032</math></td> <td><math>45.4 / 18 = 2.5222</math></td> </tr> <tr> <td>Ratio<br/>(<math>\div</math> smallest)</td> <td>1</td> <td>6.0007</td> </tr> </tbody> </table> <p>Allow <math>CoCl_2 \cdot 6H_2O</math></p> |      | $MgSO_4$ | $H_2O$ | % | 48.9 | 51.1 | Moles<br>( $\div$ RFM) | $48.9 / 120.4 = 0.406146170\dots$ | $51.1 / 18 = 2.838888889\dots$ | Ratio<br>( $\div$ smallest) | 1 | 6.98982049 |  | $CoCl_2$ | $H_2O$ | % | 54.6 | 45.4 | Moles<br>( $\div$ RFM) | $54.6 / 129.9 = 0.42032$ | $45.4 / 18 = 2.5222$ | Ratio<br>( $\div$ smallest) | 1 | 6.0007 | (4) |
|                             | $MgSO_4$   | $H_2O$   |      |          |        |   |      |      |                        |                                   |                                |                             |   |            |  |          |        |   |      |      |                        |                          |                      |                             |   |        |     |
| %                           | 48.9   | 51.1   |      |          |        |   |      |      |                        |                                   |                                |                             |   |            |  |          |        |   |      |      |                        |                          |                      |                             |   |        |     |
| Moles<br>( $\div$ RFM)      | $48.9 / 120.4 = 0.406146170\dots$  | $51.1 / 18 = 2.838888889\dots$   |      |          |        |   |      |      |                        |                                   |                                |                             |   |            |  |          |        |   |      |      |                        |                          |                      |                             |   |        |     |
| Ratio<br>( $\div$ smallest) | 1  | 6.98982049   |      |          |        |   |      |      |                        |                                   |                                |                             |   |            |  |          |        |   |      |      |                        |                          |                      |                             |   |        |     |
|                             | $CoCl_2$   | $H_2O$   |      |          |        |   |      |      |                        |                                   |                                |                             |   |            |  |          |        |   |      |      |                        |                          |                      |                             |   |        |     |
| %                           | 54.6   | 45.4   |      |          |        |   |      |      |                        |                                   |                                |                             |   |            |  |          |        |   |      |      |                        |                          |                      |                             |   |        |     |
| Moles<br>( $\div$ RFM)      | $54.6 / 129.9 = 0.42032$   | $45.4 / 18 = 2.5222$   |      |          |        |   |      |      |                        |                                   |                                |                             |   |            |  |          |        |   |      |      |                        |                          |                      |                             |   |        |     |
| Ratio<br>( $\div$ smallest) | 1  | 6.0007   |      |          |        |   |      |      |                        |                                   |                                |                             |   |            |  |          |        |   |      |      |                        |                          |                      |                             |   |        |     |



|   |     |  |  |
|---|-----|--|--|
| Method 2  |     | <u>Example of calculation:</u>   |  |
| Calculates $M_r$ of $\text{MgSO}_4$   | (1) | $M_r$ of $\text{MgSO}_4 = 24.3 + 32.1 + (4 \times 16) = 120.4$                             |  |
| Forms algebraic equation for $M_r$ of $\text{MgSO}_4 \cdot x\text{H}_2\text{O}$                               | (1) | $M_r$ of $\text{MgSO}_4 \cdot x\text{H}_2\text{O} = 120.4 + 18x$                           |  |
| Finds algebraic expression for ratio of $\text{MgSO}_4$ to hydrated $\text{MgSO}_4$                           | (1) | $\frac{120.4}{120.4 + 18x} = 48.9 \%$  |  |
| Solves for x  | (1) | $x = 7$  |  |
| For Alternative   |     |  |  |
| Calculates $M_r$ of $\text{CoCl}_2$   | (1) | $M_r$ of $\text{CoCl}_2 = 58.9 + 2 \times 35.5 = 129.9$                                    |  |
| Forms algebraic equation for $M_r$ of $\text{CoCl}_2 \cdot x\text{H}_2\text{O}$                               | (1) | $M_r$ of $\text{CoCl}_2 \cdot x\text{H}_2\text{O} = 129.9 + 18x$                           |  |
| Finds algebraic expression for ratio of $\text{CoCl}_2$ to hydrated $\text{CoCl}_2 \cdot x\text{H}_2\text{O}$ | (1) | $\frac{129.9}{129.9 + 18x} = 54.6 \%$  |  |
| Solves for x  | (1) | $x = 6$  |  |
| Method 3  |     |  |  |
| Calculates $M_r$ of $\text{MgSO}_4$   | (1) | $M_r$ of $\text{MgSO}_4 = 24.3 + 32.1 + (4 \times 16.0) = 120.4$                           |  |
| Calculates $M_r$ of $\text{MgSO}_4 \cdot x\text{H}_2\text{O}$   | (1) | $M_r$ of $\text{MgSO}_4 \cdot x\text{H}_2\text{O} = \frac{120.4 \times 100}{48.9} = 246.2$ |  |
| Calculates mass of water in one mol   | (1) | $246.2 - 120.4 = 125.8$  |  |
| Finds moles of water  | (1) | $\frac{125.8}{18} = 7$   |  |

|  |  |   |  |
|--|--|---|--|
|  | <p><u>Method 3 for <math>\text{CoCl}_2 \cdot y\text{H}_2\text{O}</math></u><br/> Calculates <math>M_r</math> of for <math>\text{CoCl}_2</math>. (1)</p> <p>Calculates <math>M_r</math> of <math>\text{CoCl}_2 \cdot y\text{H}_2\text{O}</math><br/> (1)</p> <p>Calculates mass of water in one mol (1)</p> <p>Finds moles of water (1)</p> | $M_r \text{ of } \text{CoCl}_2 = 58.9 + (2 \times 35.5) = 129.9$ $M_r \text{ of } \text{CoCl}_2 \cdot y\text{H}_2\text{O} = \frac{129.9 \times 100}{54.6} = 237.9$ $237.9 - 129.9 = 108.0$ $\frac{108.0}{18} = 6$ <p>Use of Beryllium<br/> Calculates <math>M_r</math> of <math>\text{BeSO}_4 = 105.1</math><br/> Moles in 48.9% of 100g = 0.46527117<br/> Ratio of <math>\text{BeSO}_4:\text{H}_2\text{O} = 1:6.102</math><br/> <math>x=6</math></p> |  |
|--|--|---|--|

(Total for Question 4 = 8 marks)

| Question Number | Acceptable Answer  | Additional Guidance   | Mark |
|-----------------|--|---|------|
| 5(a)(i)         | <p>A description making reference to the following points:</p> <ul style="list-style-type: none"> <li>fizzing / effervescence stops (1)</li> <li>(all) metal carbonate / solid disappears (1)</li> </ul> | <p>Allow stops frothing / no more bubbles</p> <p>Allow metal carbonate / solid "dissolved"</p> <p><b>OR just 'a clear solution forms' for M2</b></p> <p>Ignore colourless</p> | (2)  |

| Question Number | Acceptable Answer   | Additional Guidance   | Mark |
|-----------------|---|---|------|
| 5(a)(ii)        | <ul style="list-style-type: none"> <li>remove excess / unreacted metal carbonate</li> </ul> | <p>Allow to remove excess / unreacted solid</p> <p><b>Allow "removes insoluble solid"</b></p> <p>Ignore just "to remove impurities"</p> | (1)  |

| Question Number | Acceptable Answer  | Additional Guidance   | Mark |
|-----------------|--|---|------|
| 5(a)(iii)       | <p>An explanation that makes reference to the following points:</p> <ul style="list-style-type: none"> <li>so as little product dissolves as possible (1)</li> <li>to remove any soluble impurities (1)</li> </ul> | <p>Allow product might dissolve in large volumes / warm water</p> <p>Ignore rinse / wash / clean the crystals</p> <p>Ignore hydration of crystals</p> | (2)  |

| Question Number | Acceptable Answer   | Additional Guidance   | Mark |
|-----------------|---|---|------|
| 5(b)            | <ul style="list-style-type: none"> <li>• M1 calculate moles of acid (1)</li> <li>• M2 finds moles of <math>\text{JCl}_2 / 6\text{H}_2\text{O}</math> (1)</li> </ul> <p>Either</p> <ul style="list-style-type: none"> <li>• M3 finds <math>M_r</math> of <math>\text{JCl}_2</math> (1)</li> <li>• M4 finds <math>A_r</math> of J (1)</li> </ul> <p>Or</p> <ul style="list-style-type: none"> <li>• M3 finds mass of water and finds mass of <math>\text{JCl}_2</math> by subtraction (1)</li> <li>• M4 finds mass and <math>A_r</math> of J (1)</li> </ul> <p>Or</p> <ul style="list-style-type: none"> <li>• M5 identifies J (1)</li> </ul> | <p><u>Example of calculation:</u></p> <p><math>150/1000 \times 0.800 = 0.12(0)</math> (mol)</p> <p><math>0.12 / 2 = 0.06(00)</math> (mol)</p> <p>Either</p> <p><math>M_r = 14.26 / 0.0600 = 237.7</math> (<math>\text{g mol}^{-1}</math>)</p> <p><math>A_r = 237.7 - (71 + 108) = 58.7</math> (<math>\text{g mol}^{-1}</math>)</p> <p>Or</p> <p>J is Ni<br/>Allow TE for M5 on the <math>A_r</math> calculated</p> <p>Mass of water = <math>0.06 \times 6 \times 18 = 6.48</math> (g)<br/>Mass of <math>\text{JCl}_2 = 14.26 - 6.48 = 7.78</math> (g)</p> <p>Mass of J = <math>7.78 - (0.06 \times 71) = 3.52</math> (g)<br/><math>A_r</math> of J = <math>\frac{3.52}{0.06} = 58.66667 / 58.7</math> (<math>\text{g mol}^{-1}</math>)</p> <p>Or</p> <p><math>M_r</math> of <math>\text{JCl}_2 = \frac{7.78}{0.06} = 129.6667 / 129.7</math><br/><math>A_r = 129.7 - 71 = 58.7</math> (<math>\text{g mol}^{-1}</math>)<br/>J is Ni<br/>Allow TE for M5 on the <math>A_r</math> calculated</p> <p>Ignore SF except 1SF</p> | (5)  |

| Question Number | Acceptable Answer   | Additional Guidance  | Mark |
|-----------------|---|--|------|
| 5(c)(i)         | <p>An explanation which makes the following points:</p> <p>M1</p> <ul style="list-style-type: none"> <li>transition metals form coloured compounds / are not normally white</li> </ul> <p>or</p> <ul style="list-style-type: none"> <li>crystals are white suggesting (compound of) an s-block element / group 2 element (1)</li> </ul> <p>M2</p> <ul style="list-style-type: none"> <li>flame test to identify cation / metal ion (1)</li> </ul> | <p>Allow any stated colour as long as the presence of a transition metal (in the compound) is stated</p> <p>Do not award compound of a group 1 element</p> | (2)  |

| Question Number | Answer  | Mark |
|-----------------|---|------|
| 5(c)(ii)        | <p>The only correct answer is D</p> <p><i>A is not correct because barium gives a green flame colour</i></p> <p><i>B is not correct because calcium gives an orange-red flame colour</i></p> <p><i>C is not correct because lithium is not in Group 2</i></p> | (1)  |

| Question Number | Acceptable Answer   | Additional Guidance  | Mark |
|-----------------|---|--|------|
| 5(c) (iii)      | <p>Method 1<br/>calculate molar mass of <math>\text{SrCl}_2 \cdot 6\text{H}_2\text{O}</math> (1)</p> <p>EITHER</p> <p>calculates the percentage yield (1)<br/>or</p> <p>calculates maximum mass of <math>\text{SrCl}_2 \cdot 6\text{H}_2\text{O}</math> and hence percentage yield (1)<br/>or</p> <p>finds moles of <math>\text{SrCl}_2 \cdot 6\text{H}_2\text{O}</math> and hence percentage yield (1)</p> | <p><u>Example of Calculation</u><br/>Correct answer with no working scores (2)<br/>266.6</p> <p><math>237.7 / 266.6 \times 100 = 89.16\%</math><br/><math>= 89\%</math> (to 2 S.F.)</p> <p>Maximum mass = <math>0.0600 \times 266.6</math><br/><math>= 15.996</math> (g)</p> <p>Percentage yield<br/><math>= \frac{14.26}{15.996} \times 100 = 89.147\%</math><br/><math>= 89\%</math> (to 2 S.F.)</p> <p><math>\frac{14.26}{266.6} = 0.0534883 / 0.0535</math> (mol)</p> <p>Moles of <math>\text{SrCO}_3 / \text{SrCl}_2</math> (calculated in 5(b))<br/><math>= 0.06</math> (mol)</p> <p>Percentage yield<br/><math>= \frac{0.0534883}{0.0600} \times 100 = 89.147\%</math><br/><math>= 89\%</math> (to 2 S.F.)</p> <p>Allow TE on an incorrect choice of metal only</p> | (2)  |

(Total for Question 5 = 15 marks)

| Question Number | Acceptable Answer        | Additional Guidance   | Mark |
|-----------------|--------------------------|---|------|
| 6(a) (i)        | $(1s^2)2s^22p^63s^23p^5$ | Ignore repeat of $1s^2$<br>Allow <b>1s2 2s2.....</b><br><b>1S2 2S2.....</b><br>For $3p^5$ accept $3p_x^2, 3p_y^2, 3p_z^1$ | (1)  |

| Question Number | Acceptable Answer  | Additional Guidance  | Mark |
|-----------------|--|--|------|
| 6(a) (ii)       | An explanation that makes reference to the following points: <ul style="list-style-type: none"> <li>iodine (also) has 7 electrons in the outer shell / is <math>5s^25p^5</math> / is (also) <math>np^5</math> (1)</li> <li>electronic configurations / number of electrons in the outer shell govern their chemical reactions (1)</li> </ul> | Allow has the same number of electrons in the outer shell / valence electrons for M1<br><br>M2 is dependent on M1 being scored | (2)  |

| Question Number | Acceptable Answer      | Additional Guidance                      | Mark   |     |
|-----------------|------------------------|--|--|-----|
| 6(b) (i)        | Any two correct (1)    | Ion                                      | Oxidation number of sulfur                       | (2) |
|                 | Third also correct (1) | $S_2O_3^{2-}$                            | +2 / 2+ / +II / II+                              |     |
|                 |                        | $SO_4^{2-}$                              | +6 / 6+ / +VI / VI+                              |     |
|                 |                        | $S_4O_6^{2-}$                            | +2.5 / 2.5+ / $+\frac{10}{4}$ / $\frac{10}{4}$ + |     |
|                 |                        | Allow any equivalent fractions e.g. 5/2+ |  |     |
|                 |                        | Penalise missing + once only             |  |     |

| Question Number | Acceptable Answer  | Additional Guidance   | Mark |
|-----------------|--|---|------|
| 6(b)(ii)        | An answer that makes reference to: <ul style="list-style-type: none"> <li>gain of electrons (by iodine / I<sub>2</sub>)</li> </ul> | Allow thiosulfate ion has lost electrons / sulfur has lost electrons<br><br>Ignore reference to oxidation numbers | (1)  |

| Question Number | Acceptable Answer   | Additional Guidance   | Mark |
|-----------------|---|---|------|
| 6(b)(iii)       | An answer that makes reference to: <ul style="list-style-type: none"> <li>chlorine oxidises sulfur (from +2) to +6 whereas iodine only oxidises sulfur (from +2) to +2.5</li> </ul> | Allow chlorine causes a greater increase in oxidation number (than iodine)<br>OR<br>chlorine causes loss of more electrons (from sulfur than iodine)<br>Do not award chlorine gains more electrons<br><br>Award mark for a greater increase in oxidation number, even if the stated oxidation numbers are incorrect | (1)  |



| Question Number | Acceptable Answer   | Additional Guidance   | Mark |
|-----------------|---|---|------|
| 6(b)(iv)        | <ul style="list-style-type: none"> <li>correct species (1)</li> <li>balancing of correct species (1)</li> </ul> | <p><u>Example of equation</u></p> $\text{S}_2\text{O}_3^{2-} + 5\text{H}_2\text{O} \rightarrow 2\text{SO}_4^{2-} + 10\text{H}^+ + 8\text{e}^-$ <p>Allow for one mark:</p> $\text{S}_2\text{O}_3^{2-} + 10\text{OH}^- \rightarrow 2\text{SO}_4^{2-} + 5\text{H}_2\text{O} + 8\text{e}^-$ <p>Ignore state symbols even if incorrect</p> | (2)  |

| Question Number | Acceptable Answer  | Additional Guidance  | Mark |
|-----------------|--|--|------|
| 6(b)(v)         | <ul style="list-style-type: none"> <li>correct equation</li> </ul> | <p><u>Example of equation</u></p> $4\text{Cl}_2 + \text{S}_2\text{O}_3^{2-} + 5\text{H}_2\text{O} \rightarrow 8\text{Cl}^- + 2\text{SO}_4^{2-} + 10\text{H}^+$ <p>Allow HCl in place of H<sup>+</sup> and Cl<sup>-</sup> as long as balanced (8HCl + 2H<sup>+</sup>)</p> <p>Allow</p> $4\text{Cl}_2 + \text{S}_2\text{O}_3^{2-} + 10\text{OH}^- \rightarrow 8\text{Cl}^- + 2\text{SO}_4^{2-} + 5\text{H}_2\text{O}$ <p>From</p> $\text{S}_2\text{O}_3^{2-} + 10\text{OH}^- \rightarrow 2\text{SO}_4^{2-} + 5\text{H}_2\text{O} + 8\text{e}^-$ <p>in (b)(iv)</p> <p>Do not award equations with electrons not cancelled</p> <p>Ignore state symbols even if incorrect</p> | (1)  |

(Total for Question 6 = 10 marks)

| Question Number                                    | Acceptable Answer  | Additional Guidance                                | Mark  |   |   |     |   |     |   |   |   |   |   |   |     |
|--|--|--|---|---|---|-----|---|-----|---|---|---|---|---|---|-----|
| * 7  | <p><b>This question assesses a student’s ability to show a coherent and logically structured answer with linkages and fully-sustained reasoning.</b></p> <p>Marks are awarded for indicative content and for how the answer is structured and shows lines of reasoning.</p> <p>The following table shows how the marks should be awarded for indicative content.</p> <table border="1" data-bbox="383 660 1158 930"> <thead> <tr> <th data-bbox="383 660 770 759">Number of indicative marking points seen in answer</th> <th data-bbox="770 660 1158 759">Number of marks awarded for indicative marking points</th> </tr> </thead> <tbody> <tr> <td data-bbox="383 759 770 794">6</td> <td data-bbox="770 759 1158 794">4</td> </tr> <tr> <td data-bbox="383 794 770 829">5-4</td> <td data-bbox="770 794 1158 829">3</td> </tr> <tr> <td data-bbox="383 829 770 865">3-2</td> <td data-bbox="770 829 1158 865">2</td> </tr> <tr> <td data-bbox="383 865 770 900">1</td> <td data-bbox="770 865 1158 900">1</td> </tr> <tr> <td data-bbox="383 900 770 930">0</td> <td data-bbox="770 900 1158 930">0</td> </tr> </tbody> </table> | Number of indicative marking points seen in answer | Number of marks awarded for indicative marking points | 6 | 4 | 5-4 | 3 | 3-2 | 2 | 1 | 1 | 0 | 0 | <p>Guidance on how the mark scheme should be applied:</p> <p>The mark for indicative content should be added to the mark for lines of reasoning.</p> <p>For example, an answer with five indicative marking points, which is partially structured with some linkages and lines of reasoning, scores 4 marks (3 marks for indicative content and 1 mark for partial structure and some linkages and lines of reasoning).</p> <p>If there are no linkages between points, the same five indicative marking points would yield an overall score of 3 marks (3 marks for indicative content and no marks for linkages).</p> | (6) |
| Number of indicative marking points seen in answer | Number of marks awarded for indicative marking points  |  |   |   |   |     |   |     |   |   |   |   |   |   |     |
| 6  | 4  |  |   |   |   |     |   |     |   |   |   |   |   |   |     |
| 5-4  | 3  |  |   |   |   |     |   |     |   |   |   |   |   |   |     |
| 3-2  | 2  |  |   |   |   |     |   |     |   |   |   |   |   |   |     |
| 1  | 1  |  |   |   |   |     |   |     |   |   |   |   |   |   |     |
| 0  | 0  |  |   |   |   |     |   |     |   |   |   |   |   |   |     |

| Question Number   | Acceptable Answer   | Additional Guidance | Mark  |   |   |   |   |  |   |  |  |
|---|---|---------------------|---|---|---|---|---|--|---|--|--|
| *7<br>contd   | <p>The following table shows how the marks should be awarded for structure and lines of reasoning.</p> <table border="1" data-bbox="383 357 1294 751"> <thead> <tr> <th data-bbox="383 357 891 456"></th> <th data-bbox="891 357 1294 456">Number of marks awarded for structure of answer and sustained line of reasoning</th> </tr> </thead> <tbody> <tr> <td data-bbox="383 456 891 587">Answer shows a coherent and logical structure with linkages and fully sustained lines of reasoning demonstrated throughout.</td> <td data-bbox="891 456 1294 587">2</td> </tr> <tr> <td data-bbox="383 587 891 686">Answer is partially structured with some linkages and lines of reasoning.</td> <td data-bbox="891 587 1294 686">1</td> </tr> <tr> <td data-bbox="383 686 891 751">Answer has no linkages between points and is unstructured.</td> <td data-bbox="891 686 1294 751">0</td> </tr> </tbody> </table> |                     | Number of marks awarded for structure of answer and sustained line of reasoning | Answer shows a coherent and logical structure with linkages and fully sustained lines of reasoning demonstrated throughout. | 2 | Answer is partially structured with some linkages and lines of reasoning. | 1 | Answer has no linkages between points and is unstructured. | 0 | <p>In general it would be expected that 5 or 6 indicative points would get 2 reasoning marks, and 3 or 4 indicative points would get 1 mark for reasoning, and 0, 1 or 2 indicative points would score zero marks for reasoning.</p> <p>If there is any incorrect chemistry, deduct mark(s) from the reasoning. If no reasoning mark(s) awarded do not deduct mark(s).</p> |  |
|   | Number of marks awarded for structure of answer and sustained line of reasoning   |                     |   |   |   |   |   |  |   |  |  |
| Answer shows a coherent and logical structure with linkages and fully sustained lines of reasoning demonstrated throughout. | 2   |                     |   |   |   |   |   |  |   |  |  |
| Answer is partially structured with some linkages and lines of reasoning.   | 1   |                     |   |   |   |   |   |  |   |  |  |
| Answer has no linkages between points and is unstructured.  | 0   |                     |   |   |   |   |   |  |   |  |  |

|  |  |  |  |
|--|--|--|--|
|  | <p>Indicative content:</p> <ul style="list-style-type: none"> <li>• <u>IP1 Electrons</u><br/>Same number of electrons so similar / the same <b>London forces / van der Waals' forces</b> / dispersion forces</li> <li>• <u>IP2 Electronegativity</u><br/>Large electronegativity differences in HF and H<sub>2</sub>O and small in CH<sub>4</sub> / quoting all electronegativity values of differences / combination of previous three alternatives covering all three bonds</li> <li>• <u>IP3 Intermolecular forces in methane</u><br/>Only (weak) London forces / <b>van der Waals'</b> forces / dispersion forces in CH<sub>4</sub></li> <li>• <u>IP4 Intermolecular forces in water and hydrogen fluoride</u><br/>Hydrogen bonding in both HF and H<sub>2</sub>O (but not CH<sub>4</sub>)</li> <li>• <u>IP5 Relative numbers of hydrogen bonds</u><br/>More hydrogen bonds / (average of) twice as many hydrogen bonds in H<sub>2</sub>O than in HF</li> <li>• <u>IP6 Energy</u><br/>More energy needed to break stronger intermolecular forces / less needed to break weaker intermolecular forces.</li> </ul> | <p>Read all of the answer first as IPs can be found anywhere in the answer</p> <p>Allow high electronegativity of F and O (compared to H)<br/>Allow HF and H<sub>2</sub>O (highly) polar and CH<sub>4</sub> non polar</p> <p>Allow IP2 for any three of:<br/>F=4.0, O=3.5, H=2.1, C=2.5<br/>Allow IP2 for any two of:<br/>HF=1.9, HO=1.4, HC=0.4<br/>These values may be seen anywhere</p> <p>Allow no dipole-dipole forces / no hydrogen bonds in CH<sub>4</sub><br/>Award IP3 if London forces are the only intermolecular forces mentioned in CH<sub>4</sub></p> <p>May be shown in a diagram</p> <p>Do not award IP6 for any clear indication of covalent bond breaking or ionic bond breaking</p> |  |
|--|--|--|--|

(Total for Question 7 = 6 marks)

| Question Number | Answer  | Mark |
|-----------------|---|------|
| 8(a)(i)         | The only correct answer is B<br><i>A is not correct because fluorine is diatomic</i><br><i>C is not correct because sodium is 1<sup>+</sup> ion</i><br><i>D is not correct because fluorine is diatomic</i> | (1)  |



| Question Number | Answer   | Mark |
|-----------------|--|------|
| 8(a) (iii)      | <p>The only correct answer is A</p> <p><i>B is not correct because diagram has cations larger than anions</i></p> <p><i>C is not correct because diagram has cations larger than anions</i></p> <p><i>D is not correct because trends in wrong direction</i></p> | (1)  |

| Question Number | Acceptable Answer  | Additional Guidance  | Mark |
|-----------------|--|--|------|
| 8(a) (iv)       | <ul style="list-style-type: none"> <li>increase in number of protons (in the nucleus) (1)</li> <li>increases the attraction for the electrons (bringing them closer to the nucleus) (1)</li> </ul> | <p>Allow increasing nuclear charge</p> <p>For explanations of graph B allow max (1) for a correct explanation for any downward trend for three ions</p> <p>Allow max (1) for an explanation of the smallest or largest ion without an explanation of the trend</p> <p>e.g. <math>\text{Al}^{3+}</math> has the most protons so electrons most attracted to nucleus so smallest scores (1)</p> <p>Discussion of atomic radius max (1)</p> | (2)  |

| Question Number | Acceptable Answer  | Additional Guidance   | Mark |
|-----------------|--|---|------|
| 8(b)            | <p>An explanation that makes reference to the following points:</p> <ul style="list-style-type: none"> <li>the higher the charge on the cation the stronger the attraction between ions and mention of a 2+ cation in CaF<sub>2</sub> compared to a 1+ cation in LiF / KF (1)</li> <li>the smaller the radius of the cation the stronger the attraction between ions and mention of Li<sup>+</sup> being smaller than K<sup>+</sup> (1)</li> </ul> | <p><b>Allow "stronger bonding" for stronger attraction between ions</b></p> <p>Both charges should be stated<br/>Allow calcium ions have twice the charge of potassium / lithium ions.</p> <p><b>Do not award 'lithium has a smaller radius than potassium' unless it is clear ions are being considered, for example the use of Li<sup>+</sup> and K<sup>+</sup> in the answer.</b></p> <p>If no other marks awarded, allow a discussion of charge density without reference to charge or radius of one pair of ions for (1)</p> <p>If no other mark awarded, allow a correct statement about the effect of charge and ionic radius without justification from table of data for (1)</p> | (2)  |

(Total for Question 8 = 9 marks)

TOTAL FOR PAPER = 80 MARKS



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